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(Remember—you must base this answer on how many moles of the limiting reactant that reacted.) 1.5 mol. 0.75 mol of MgCl_2 is the limiting reactant, so we use the 6:3 ratio of NaCl to MgCl_2 and multiply $(0.75)(6/3) = 1.5$ mol.

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the number of moles of O_2 needed are $(12.5)(5) = 62.5$ moles.

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100 ChemQuest 33 Name: _____ Date: _____ Hour: _____

Information: Limiting Reactant Again consider the combustion of propane: $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$. If you had 10 moles of propane to burn, you would need 50 moles of oxygen according to the ratio in the balanced equation. If you only had 20 moles of oxygen you could not combust all 10 moles of propane.

ChemQuest33 Key - 100 ChemQuest 33 Name Date Hour

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of O₂ needed are $(12.5)(5) = 62.5$ moles.

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Chemquest 33 Answers - vpn.sigecloud.com.br Chemquest 33 Limiting Reactants Answers Verify that Na₃PO₄ is the excess reactant and MgCl₂ is the limiting reactant. Using the 3:2 ratio, we find that for 0.75 mol of MgCl₂, we need $(0.75)(2/3) = 0.5$ moles of Na₃PO₄. We have more than that and therefore

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reaction quotient was 2.1. Predict what will happen to the concentration of products and reactants as the reaction progresses. 2.1 must decrease to 0.45 by decreasing the concentration of products and increasing the concentration of reactants. 8. Given your answers to questions 6 and 7, complete the following sentences.

Critical Thinking Questions

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Chemquest 18 Answer - 0900taxiservice.nl is a measure of how much an atom attracts an electron. The higher the electronegativity, the greater the atom's attraction for electrons. Atoms that become negative ions have a much greater electronegativity than atoms that become positive ions. Chemquest 18 ...

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Answers to Worksheet #14 Limiting Reagents
A Limiting Reagent is the reactant that is completely used up in a reaction. This reagent is the one that determines the amount of product formed. Limiting reagent calculations are performed in the same manner as the stoichiometric equations on Worksheet #11.

However, with a limiting *Page 8/10*

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Limiting Reagents - Ms. Mogck's Classroom

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JJ ChemQuest 33 Name T tlt'cym ,uTg.eTlnT F.uT'r \$\$ Date Hour Information: Determining if a Bond is Polar In general the greater the difference in electronegativity between t\i/o bonding atoms, the greater the polarity of the bond.A general rule of thumb is that if the difference in ele onegativity is less than 0.5 then the bond is considered nonpolar.Ifthe difference is greater

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